

Molarity pHet

Pre-Lab

1. If you were to dissolve salt or sugar into a beaker of water, how would you know if it was a saturated solution?

See if there is any solid on the bottom.
If not, add more solute and see if it dissolves.

Please go to: https://phet.colorado.edu/sims/html/molarity/latest/molarity_en.html

- 1) Click on the "5" to run the simulation;
Click on show values to display all numeric values and explore the simulation.
- 2) For each solute determine if the solution is saturated or unsaturated at a 1 Molar solution

Solute	Saturated or unsaturated
Drink mix	unsaturated
Cobalt (II) nitrate	unsaturated
Cobalt (II) chloride	unsaturated
Potassium dichromate	saturated *
Gold (III) chloride	unsaturated
Potassium chromate	unsaturated
Nickel (II) Chloride	unsaturated
Copper (II) Sulfate	unsaturated
Potassium permanganate	saturated *

- 3) If you change the solute amount but keep the solution volume the same what happens to the Molarity?

molarity changes.
Direct relationship

- a. What is the relationship between solute amount and Molarity?

Direct

- 4) If you change the solution volume but keep the solute amount the same what happens to the Molarity?

molarity changes

- a. What is the relationship between solution amount and Molarity?

Inverse